

**T04D02 - Lewis Dot Structures & NAS Rule**

Draw the following Lewis Dot structures for molecules or ions that may or may not follow the octet rule.

Use the formula  $N - A = S$  and if it has an ***even number of shared electron*** then

- it follows the octet rule unless .....  
it is group IIA (Be) or IIIA (B) or gets more than 8 electrons around it such as in  $\text{SCl}_6$ , etc.

If it has an ***odd number of shared electrons*** then

- it does not follow the octet rule and the central atom will receive the odd electron

